

## Mole Calculation Practice Worksheet

*Answer the following questions:*

- 1) How many moles are in 25.0 grams of water?
- 2) How many grams are in 4.500 moles of  $\text{Li}_2\text{O}$ ?
- 3) How many molecules are in 23.0 moles of oxygen?
- 4) How many moles are in  $3.4 \times 10^{23}$  molecules of  $\text{H}_2\text{SO}_4$ ?
- 5) How many molecules are in 25.0 grams of  $\text{NH}_3$ ?
- 6) How many grams are in  $8.200 \times 10^{22}$  molecules of  $\text{N}_2\text{I}_6$ ?

## Mole Calculation Practice Worksheet Solutions

Answer the following questions:

- 1) How many moles are in 25.0 grams of water?

**1.39 moles**

**1 mole H<sub>2</sub>O = 18.0 g H<sub>2</sub>O**

$$\frac{25 \text{ g H}_2\text{O}}{18.0 \text{ g H}_2\text{O}} \times \frac{1 \text{ mol H}_2\text{O}}{1 \text{ mol H}_2\text{O}} = 1.39 \text{ mol H}_2\text{O}$$

- 2) How many grams are in 4.500 moles of Li<sub>2</sub>O?

**134.6 grams**

**1 mole Li<sub>2</sub>O = 29.90 g Li<sub>2</sub>O**

$$\frac{4.500 \text{ mol Li}_2\text{O}}{1 \text{ mol Li}_2\text{O}} \times \frac{29.90 \text{ g Li}_2\text{O}}{1 \text{ mol Li}_2\text{O}} = 134.6 \text{ g Li}_2\text{O}$$

- 3) How many molecules are in 23.0 moles of oxygen?

**1.38 x 10<sup>25</sup> molecules**

**1 mole oxygen molecules = 6.02 x 10<sup>23</sup> oxygen molecules**

$$\frac{23.0 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{6.02 \times 10^{23} \text{ O}_2 \text{ molecules}}{1 \text{ mol O}_2} = 1.38 \times 10^{25} \text{ O}_2 \text{ molecules}$$

- 4) How many moles are in 3.4 x 10<sup>23</sup> molecules of H<sub>2</sub>SO<sub>4</sub>?

**0.56 moles**

**1 mole anything = 6.02 x 10<sup>23</sup> anything**

$$\frac{3.4 \times 10^{23} \text{ molecules H}_2\text{SO}_4}{6.02 \times 10^{23} \text{ molecules H}_2\text{SO}_4} \times \frac{1 \text{ mol H}_2\text{SO}_4}{1 \text{ mol H}_2\text{SO}_4} = 0.56 \text{ mol H}_2\text{SO}_4$$

5) How many molecules are in 25.0 grams of  $\text{NH}_3$ ?

**$8.85 \times 10^{23}$  molecules**

**1 mole  $\text{NH}_3 = 17.0 \text{ g NH}_3$**

**1 mole anything =  $6.02 \times 10^{23}$  anything**

$$\frac{25.0 \text{ g NH}_3 \left| \begin{array}{c} 1 \text{ mol NH}_3 \\ 17.0 \text{ g NH}_3 \end{array} \right| \begin{array}{c} 6.02 \times 10^{23} \text{ molecules NH}_3 \\ 1 \text{ mol NH}_3 \end{array}}{\hspace{10em}} = 8.85 \times 10^{23} \text{ molecules NH}_3$$

6) How many grams are in  $8.200 \times 10^{22}$  molecules of  $\text{N}_2\text{I}_6$ ?

**107.5 grams**

**1 mole  $\text{N}_2\text{I}_6 = 789.4 \text{ g N}_2\text{I}_6$**

**1 mole anything =  $6.02 \times 10^{23}$  anything**

$$\frac{8.200 \times 10^{22} \text{ molecules N}_2\text{I}_6 \left| \begin{array}{c} 1 \text{ mol N}_2\text{I}_6 \\ 6.02 \times 10^{23} \text{ molecules N}_2\text{I}_6 \end{array} \right| \begin{array}{c} 789.4 \text{ g N}_2\text{I}_6 \\ 1 \text{ mol N}_2\text{I}_6 \end{array}}{\hspace{10em}} = 107.5 \text{ g N}_2\text{I}_6$$